Extra Practice for Quiz!
(Not required - just extra)

Stoichiometry Practice: Mixed Conversions

1) Translate: Aluminum is added to a copper (I) chloride solution, aluminum chloride and copper are produced.

A) 3.5 grams of aluminum would result in how many grams of copper being produced?

B) If 1.25 moles of aluminum chloride were produced, how many grams of copper (I) chloride solution were required?

2) Translate: Magnesium ribbon reacts with oxygen to produce magnesium oxide.

A) If 3.5 grams of Magnesium ribbon were used, how many moles of magnesium oxide were produced?

B) Assuming 0.35 moles of magnesium oxide were produced, how many moles of Oxygen were required?

3) Translate: Potassium Chloride and Sodium Carbonate react in a double replacement reaction.

A) 5.5 grams of potassium chloride are used, how many moles of potassium carbonate is created?

B) If 18 grams of sodium chloride are produced, how many grams of sodium carbonate were required?

4) The complete combustion of tetracarbon decahydride in oxygen.

A) 6.0 moles of oxygen will produce how many moles of each of the products?

B) 135.0 grams of tetracarbon decahydride will produce how many moles of carbon dioxide

5) The reaction of $AlBr_3$ with $Mg(OH)_2$

A) 1.05 moles of aluminum bromide produces how many moles of magnesium bromide

B) 2.75 moles of magnesium hydroxide would yield how many grams of aluminum hydroxide

6) The decomposition of hydrogen peroxide (H_2O_2) to form water and oxygen.

A) 25.0 grams of oxygen gas would be produced from how many grams of hydrogen peroxide?

B) 2.5 moles of hydrogen peroxide would produce how many moles of oxygen gas?

7) The reaction of nitric acid with potassium hydroxide to form postassium nitrate and water.

A) 35.0 grams of nitric acid would produce how many moles of water?

B) 70.0 grams of potassium hydroxide will produce how many grams of potassium nitrate?