Ca As Br

1. There elements are all in the same…. (period or group)
2. Which will have the lowest electronegativity?
3. Which will have the lowest ionization energy?
4. Which element will have the greatest nuclear pull (force)?
5. Would Calcium’s ion be larger or smaller than the original atom?
6. Which trend describes the force of attraction on a different atom’s electrons?

O Se S

1. Which element has the greatest electronegativity?
2. Which element has the highest ionization energy?
3. Smallest atomic radius?
4. Which is/are metalloids?
5. Would oxygen’s ion be larger or smaller than the neutral atom?
6. Which of the elements would have the lowest electronegativity?

N P As

1. Rank these from smallest to largest atom.
2. Which element would have the highest electronegativity?
3. Which of these elements has the most energy levels?
4. Which of these elements would have the lowest ionization energy?

N F O

1. Rank these from lowest to highest ionization energy.
2. Which of the **atoms** would be the largest?
3. Which of these **ions**would be the largest?
4. Explain why Fluorine has a smaller atomic radius than Lithium.

For the following pairs of atoms, which element has the stronger (greater) electronegativity?

1. Mg Na
2. C N
3. S Se
4. O N

For the following pairs of atoms, which element has the larger atomic radius?

1. Mg Sr
2. O N
3. S O
4. O F
5. Mn Ti

For the following pairs of atoms, which element has the higher ionization energy?

1. Na K
2. Si Ge
3. Cl F
4. O S
5. Cu V